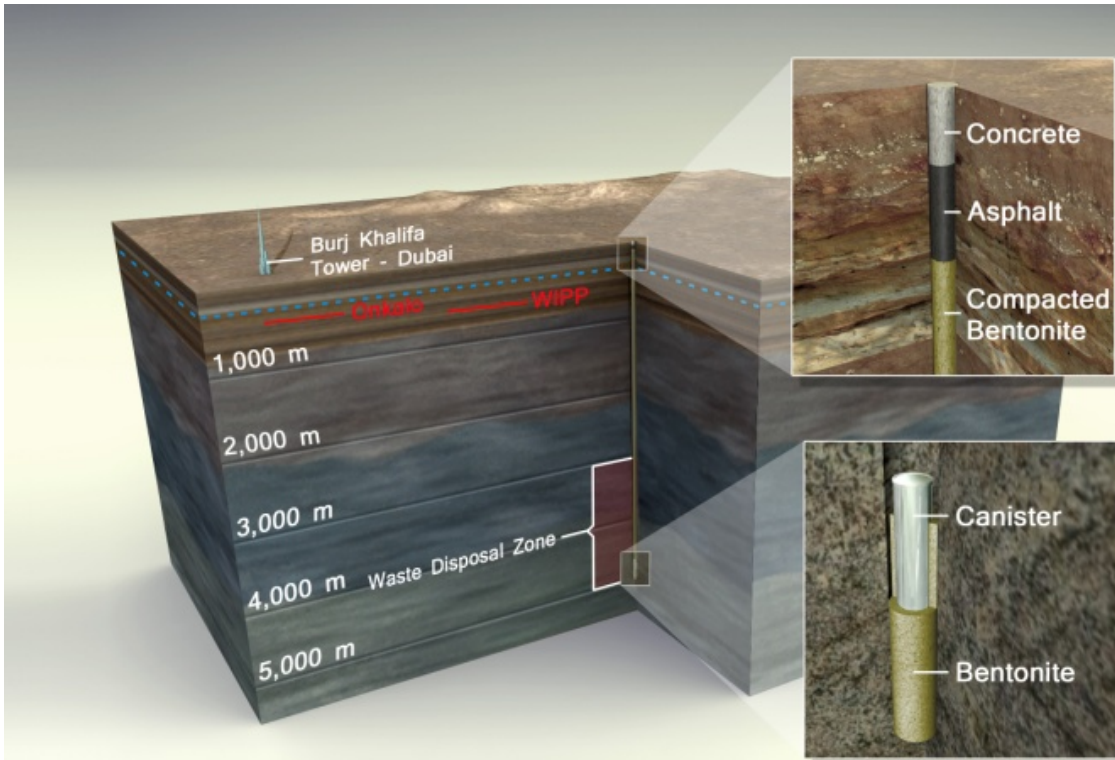


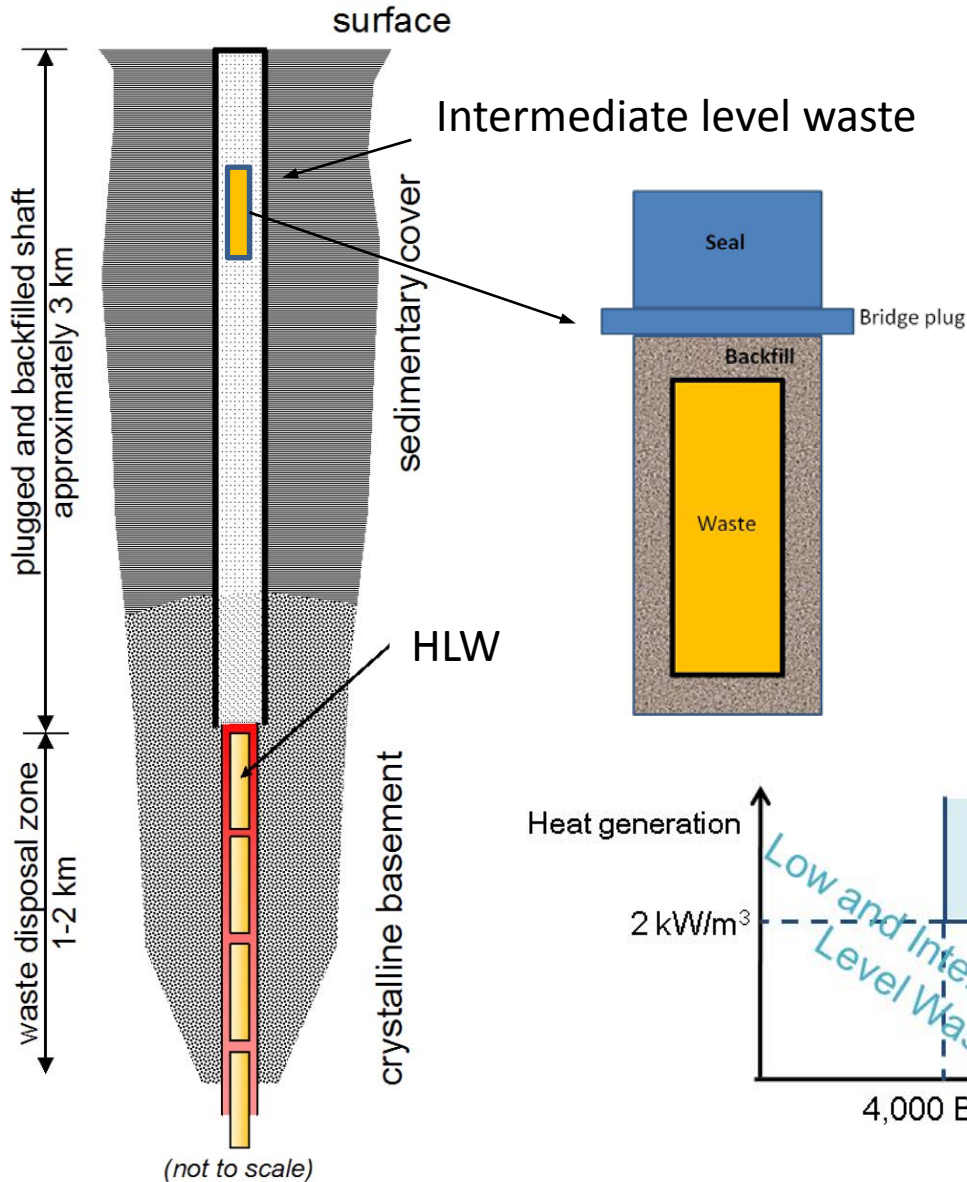
# Shallow Borehole Disposal of Intermediate Level Nuclear Waste

Patrick V. Brady, Bill Arnold, Jim Krumhansl, Courtney Herrick; Sandia National Laboratories, Albuquerque, New Mexico, pvbrady@sandia.gov



- Borehole Disposal: HLW vs. Intermediate
- 3D Thermal-hydrologic modeling
- Radionuclide solubilities/sorption
- Tailored backfills
- Seals
- Human Intrusion

# Shallow Disposal of Intermediate Level Waste

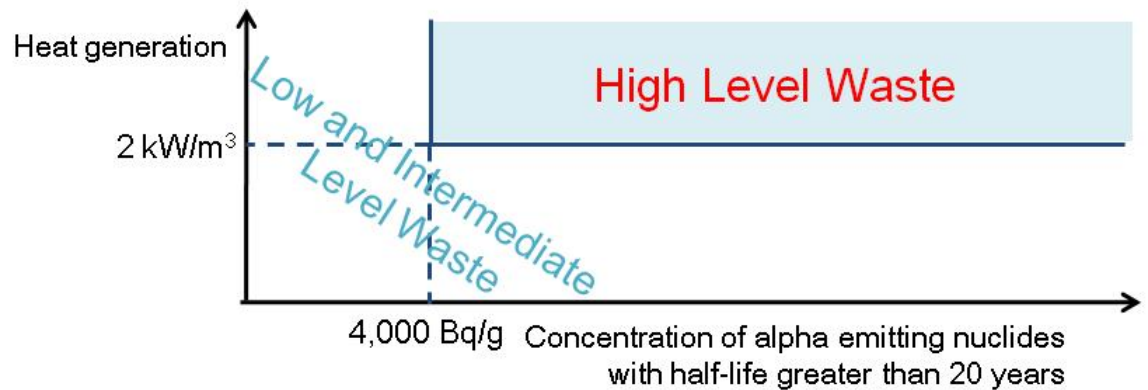


## Key Design Features

- 10s to 100s meter depth
- Metallic canister
- Multiple seals
- Tailored backfill

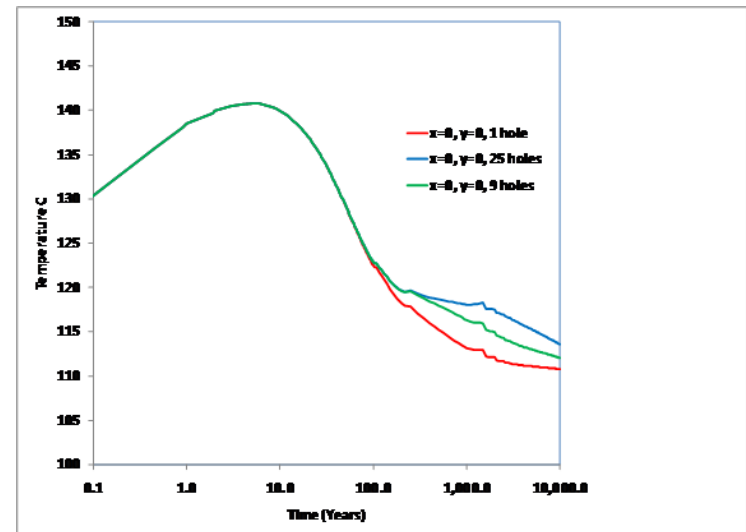
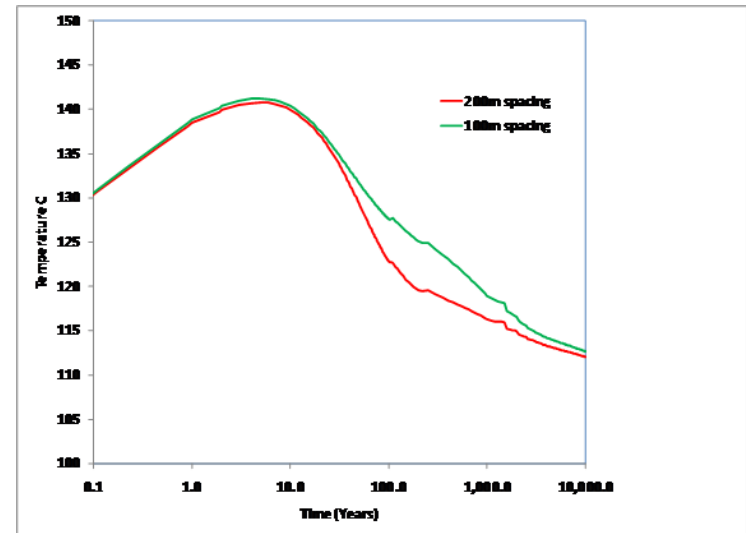
## Key Geochemical Features

- Reducing environment
- Ample bicarbonate
- Possible sulfate, sulfide



# 3D Thermal-Hydrologic Model: Multiple Boreholes

- Simulated peak temperature occurs relatively soon after waste emplacement and is insensitive to borehole spacing and number of boreholes
- Differences in temperature histories for multiple boreholes and closer borehole spacing are generally small
- Pressure and temperature conditions are below boiling for disposal of used fuel assemblies and high heat output vitrified waste from reprocessing
- Temperature perturbations are generally limited to the waste disposal zone (no significant vertical heat transport)



# Solubility and Sorption of Key Radionuclides

## Insoluble Radionuclides

Element	Solid Phase	Concentration (mol/L)	
		<sup>a</sup> Borehole	<sup>b</sup> Low E <sub>H</sub> Groundwater
Am	AmOCl	2 x 10 <sup>-8</sup>	2 x 10 <sup>-9</sup> to 10 <sup>-5</sup>
	Am <sub>2</sub> O <sub>3</sub>	3 x 10 <sup>-7</sup>	
	Am(OH) <sub>3</sub>	1 x 10 <sup>-5</sup>	
Np	Np(OH) <sub>4,am</sub>	1 x 10 <sup>-8</sup>	2 x 10 <sup>-12</sup> to 10 <sup>-8</sup>
	NpO <sub>2</sub>	2 x 10 <sup>-16</sup>	
Pu	Pu(OH) <sub>4,am</sub>	6 x 10 <sup>-6</sup>	10 <sup>-11</sup> to 10 <sup>-6</sup>
	PuO <sub>2</sub>	2 x 10 <sup>-13</sup>	
Tc	TcO <sub>2</sub> ·nH <sub>2</sub> O <sub>am</sub>	3 x 10 <sup>-8</sup>	10 <sup>-12</sup> to 10 <sup>-7</sup>
	<sup>c</sup> TcO <sub>2</sub>	9 x 10 <sup>-13</sup>	
	Tc <sub>3</sub> O <sub>4</sub>	2 x 10 <sup>-15</sup>	
	Tc sulfides	< 10 <sup>-20</sup>	
Th	Th(OH) <sub>4,am</sub>	6 x 10 <sup>-8</sup>	10 <sup>-10</sup> to 10 <sup>-9</sup>
	ThO <sub>2</sub>	4 x 10 <sup>-15</sup>	
U	<sup>d</sup> UO <sub>2,am</sub>	4 x 10 <sup>-4</sup>	10 <sup>-10</sup> to 10 <sup>-6</sup>
	UO <sub>2</sub>	6 x 10 <sup>-9</sup>	

<sup>a</sup>Borehole solubilities calculated for 150°C, 1M NaCl, C<sub>total</sub> = 1 mM, S<sub>total</sub> = 100 μmol, redox set by FeO-Magnetite equilibria. <sup>b</sup>From McKinley and Savage (1996). <sup>c</sup>25°C value from data0.ymp.r5d. All other thermodynamic values are from thermo.com.V8.R6.230. <sup>d</sup>ΔH set to -77.9 (Uraninite value).

higher salinities will  
lower values

## Soluble Radionuclides

Potential Solids	
<sup>14</sup> C	CaCO <sub>3</sub>
<sup>135,137</sup> Cs	none
<sup>129</sup> I	(Metal iodides)
<sup>226,228</sup> Ra	(SS-RaCO <sub>3</sub> , RaSO <sub>4</sub> )
<sup>90</sup> Sr	(SS-SrCO <sub>3</sub> , SrSO <sub>4</sub> )

long shots

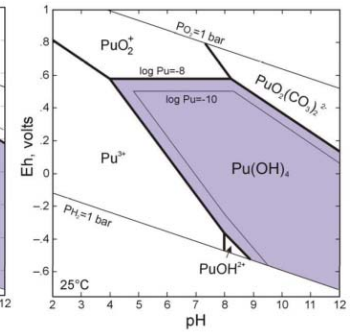
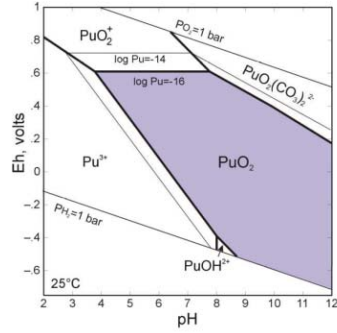
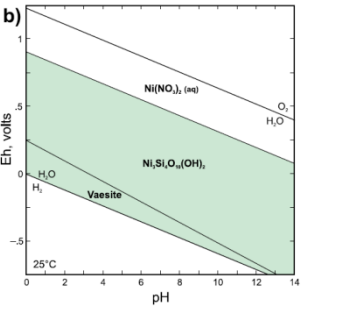
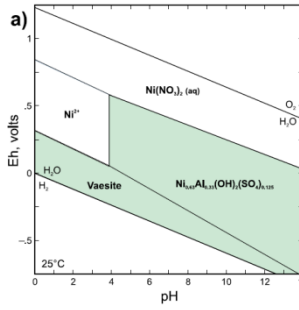
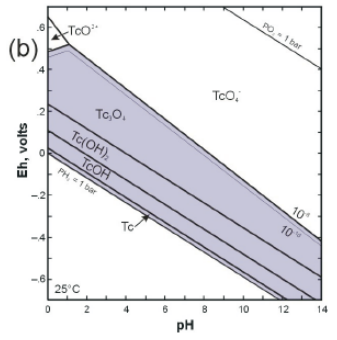
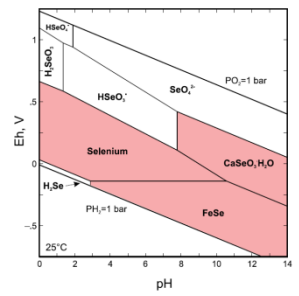
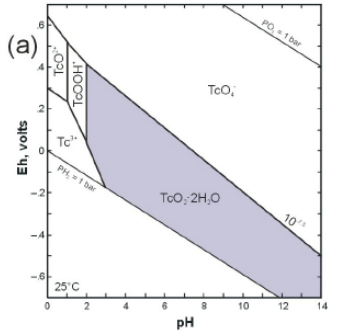
## Sorption k<sub>d</sub>s

Element	k <sub>d</sub> basement	k <sub>d</sub> sediment	k <sub>d</sub> bentonite
Am	50-5000	100-100,000	300-29,400
C	0-6	0-2000	5
Cs	50-400	10-10,000	120-1000
Np	10-5000	10-1000	30-1000
Pu	10-5000	300-100,000	150-16,800
<sup>b</sup> Ra	4-30	5-3000	50-3000
Sr	4-30	5-3000	50-3000
<sup>c</sup> Tc	0-250	0-1000	0-250
Th	30-5000	800-60,000	63-23,500
U	4-5000	20-1700	90-1000
I	0-1	0-100	0-13

<sup>a</sup>All values are from the review of McKinley and Scholtis (1993). Values less than one were rounded down to zero. <sup>b</sup>k<sub>d</sub>s for Ra were set equal to those of somewhat chemically similar Sr. <sup>c</sup>Tc k<sub>d</sub>s under reducing borehole conditions will likely be much greater than the zero values listed here which were measured under more oxidizing conditions.

# Fate of Fission Products in Shallow Borehole Disposal

Isotope	Chemical Control over Transport
$^{129}\text{I}$	Sorption to Bi-doped bentonite
$^{14}\text{C}$	Calcite formation
$^{36}\text{Cl}$	Isotopic dilution
$^{59}\text{Ni}$	Formation of Ni-Silicates, $\text{NiCO}_3(?)$ , $\text{NiS}_2$
$^3\text{H}$	*
$^{99}\text{Tc}$	Formation of $\text{TcO}_2$ , $\text{TcS}_2$
$^{135,137}\text{Cs}$	Inner layer exchange onto bentonite
$^{126}\text{Sn}$	Formation of $\text{SnO}_2$
$^{79}\text{Se}$	Formation of $\text{SeS}_2$ , $\text{FeSe}$
$^{107}\text{Pd}$	Formation of $\text{PdO-Pd(OH)}_2$



$E_H$ -pH diagrams are from Vol. 2 and 3 of USEPA's Monitored Natural Attenuation Guidelines for Metals and Radionuclides (see <http://www.epa.gov/nrmrl/pubs/600r10093/600r10093.pdf> and <http://www.epa.gov/nrmrl/pubs/600R07140/600R07140.pdf>)



# Composite Borehole Sealing

